

Part II (HONS)UNIT - 2 First Transition SeriesAssignment

1. What are d-block or transition elements?
2. Why Cu, Ag and Au are considered as transition elements but Zn, Cd, Hg are not considered as transition elements?
3. How d-block are divided into series? Write the series of d-block elements.
4. Why transition elements show metallic character?
5. Give reasons.
  - a) Size of transition elements decreases with increase in atomic no.
  - b) Transition elements have high density and ionisation energy.
  - c) Transition elements show variable valency or oxidation state.
  - d) Ionisation energy of transition elements lies between s-block and p-block elements.
  - e) Transition elements have complex formation tendency.
  - f) Transition elements have colour.
  - g) Transition elements and their compounds have catalytic property.