



Langat Singh College, Muzaffarpur  
NAAC Grade 'A'  
Under B. R. A. Bihar University, Muzaffarpur

# Mean Free Path Lec -01

**Dr. Tarun Kumar Dey**

**Professor in Physics**

**HOD, Electronics**

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**Mean free path** : The mean free path is average distance that a molecule travels between successive collisions .

If  $\lambda_1, \lambda_2, \lambda_3, \dots, \lambda_n$  are the values of successive free paths , then the total time in which they are travelled =  $\lambda_1 + \lambda_2 + \lambda_3 + \dots + \lambda_n = vt$

Where  $v$  is the mean speed of a molecule and  $n$  is the number of collisions suffered or free path traversed in time  $t$  .

If  $\lambda$  is the mean free path then

$$\lambda = \lambda_1 + \lambda_2 + \lambda_3 + \dots + \lambda_n / n = vt / n = S / n$$

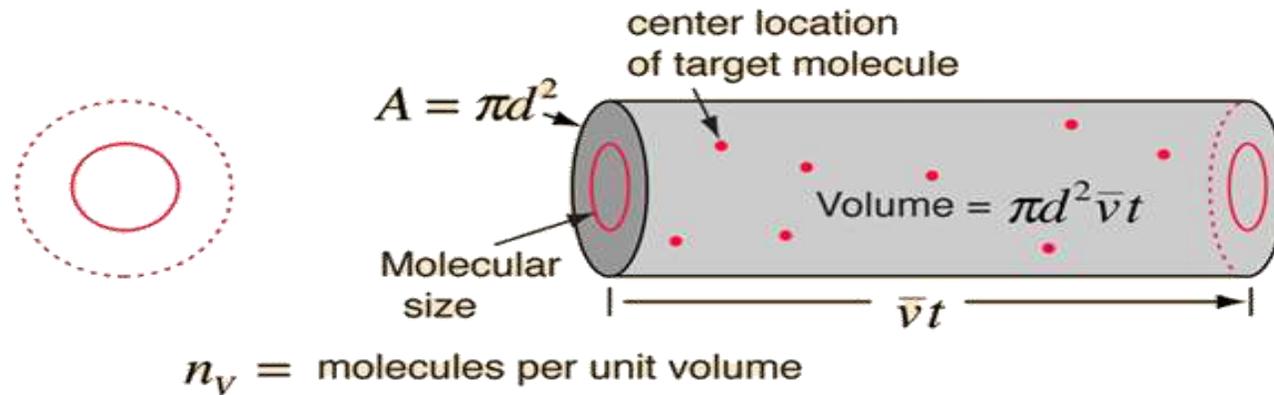
Where  $s$  is the total distance travelled in  $n$  collisions

**Expression of mean free path** : Let us assume that except one all other molecules are at rest

Then ,

$$\lambda = v / \pi d^2 v n = 1 / \pi d^2 n$$

Where  $d$  = diameter of molecules .whose centres are at a distance  $d$  from its centre .



## Clausius expression for mean free path :

According to Clausius consideration , all the molecules were moving with a constant speed in all possible direction obtained an expression

$$\lambda = \frac{3}{4\pi d^2 n} \quad (1)$$

**Maxwell's expression** : According to Maxwell consideration all the molecules are moving in all possible directions and with all possible velocities as a result of Maxwell velocity distribution law of velocities . The mean free path averaged over molecules of all velocities is given by by

$$\lambda = 1 / \pi d^2 n ( c/r) \quad (2)$$

Where  $c$  = average velocity of molecule

$r$  = mean relative velocity of all molecules w.r.t all others .