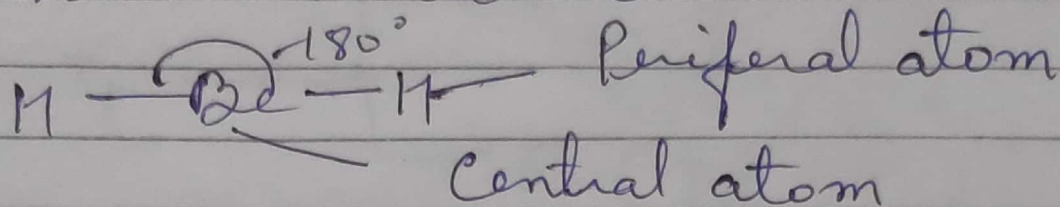


M.O diagram of BeH_2

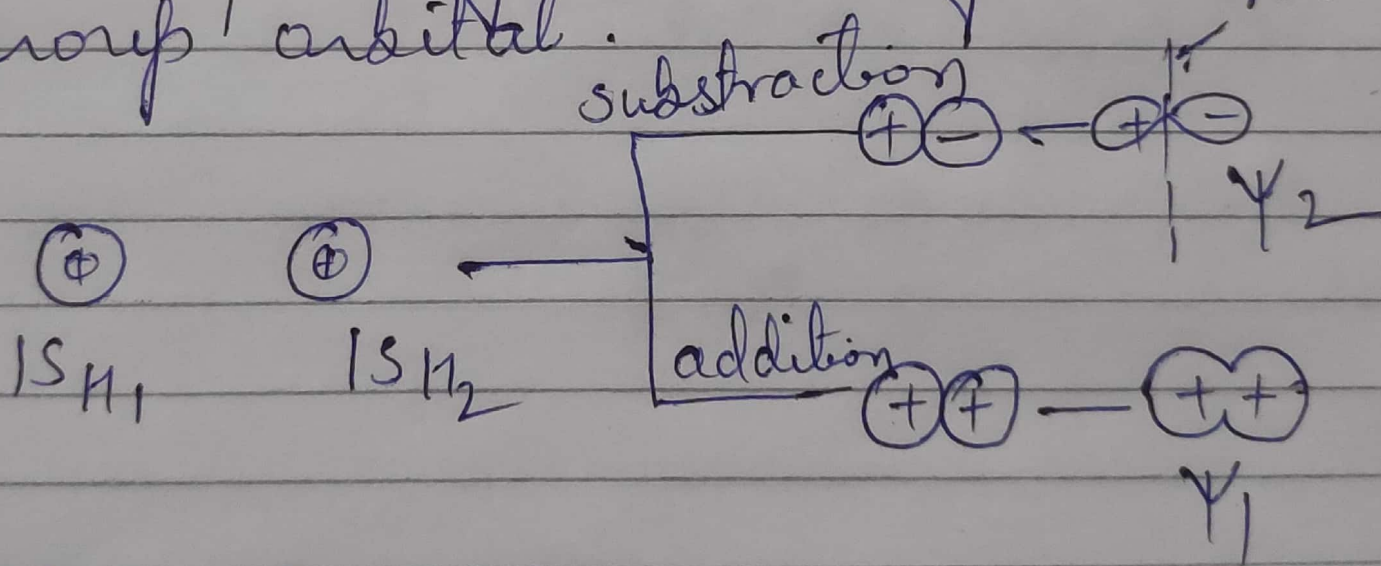
Beryllium dihydride is a linear triatomic molecule



Electronic configuration of Be is $1s^2 2s^2 2p^0$

Electronic config of H atom is $1s^1$

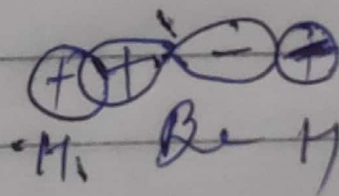
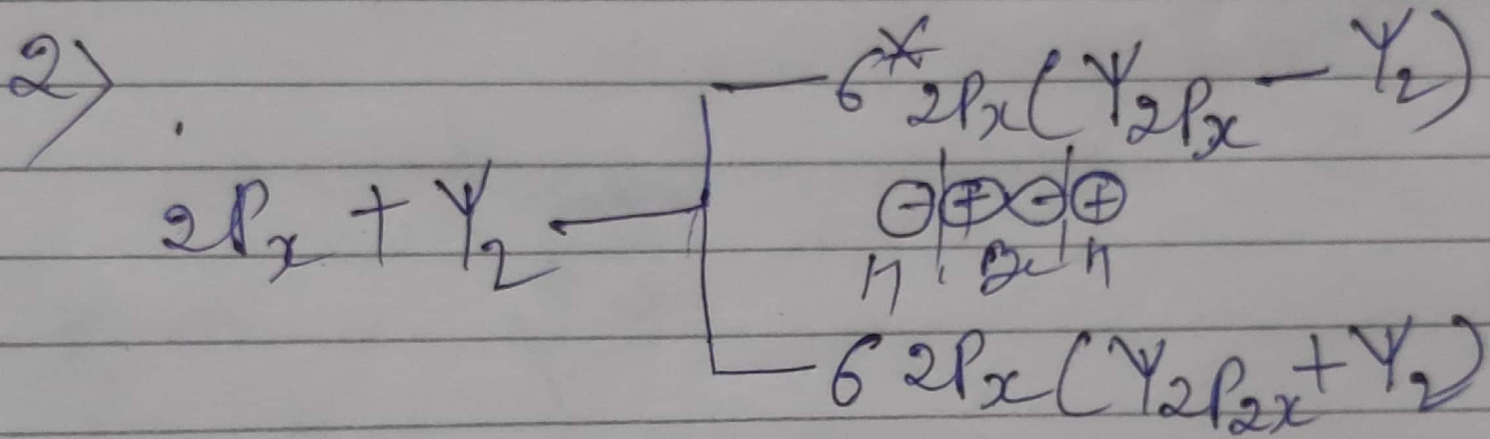
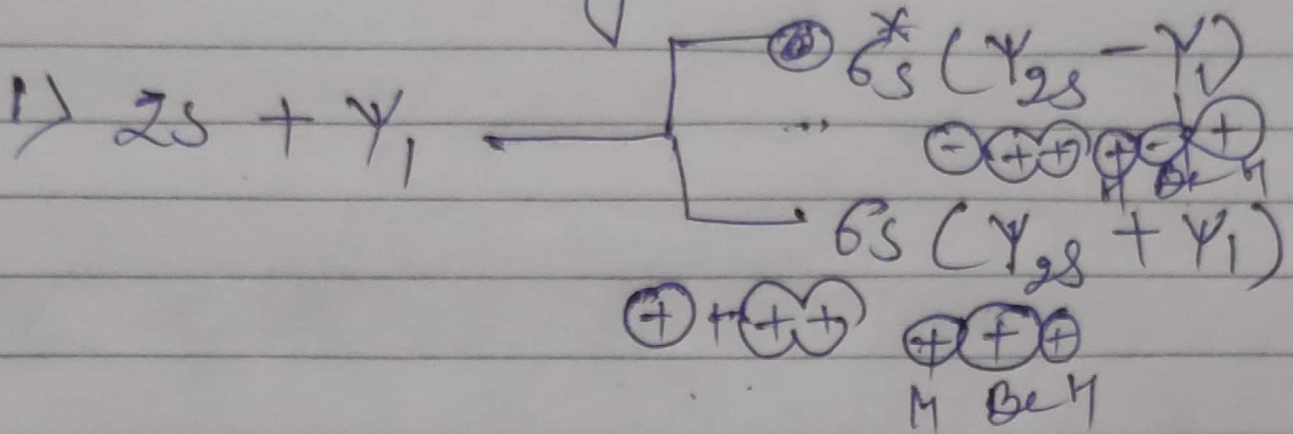
$1s$ orbital of two hydrogen atom overlap together to form group orbital.

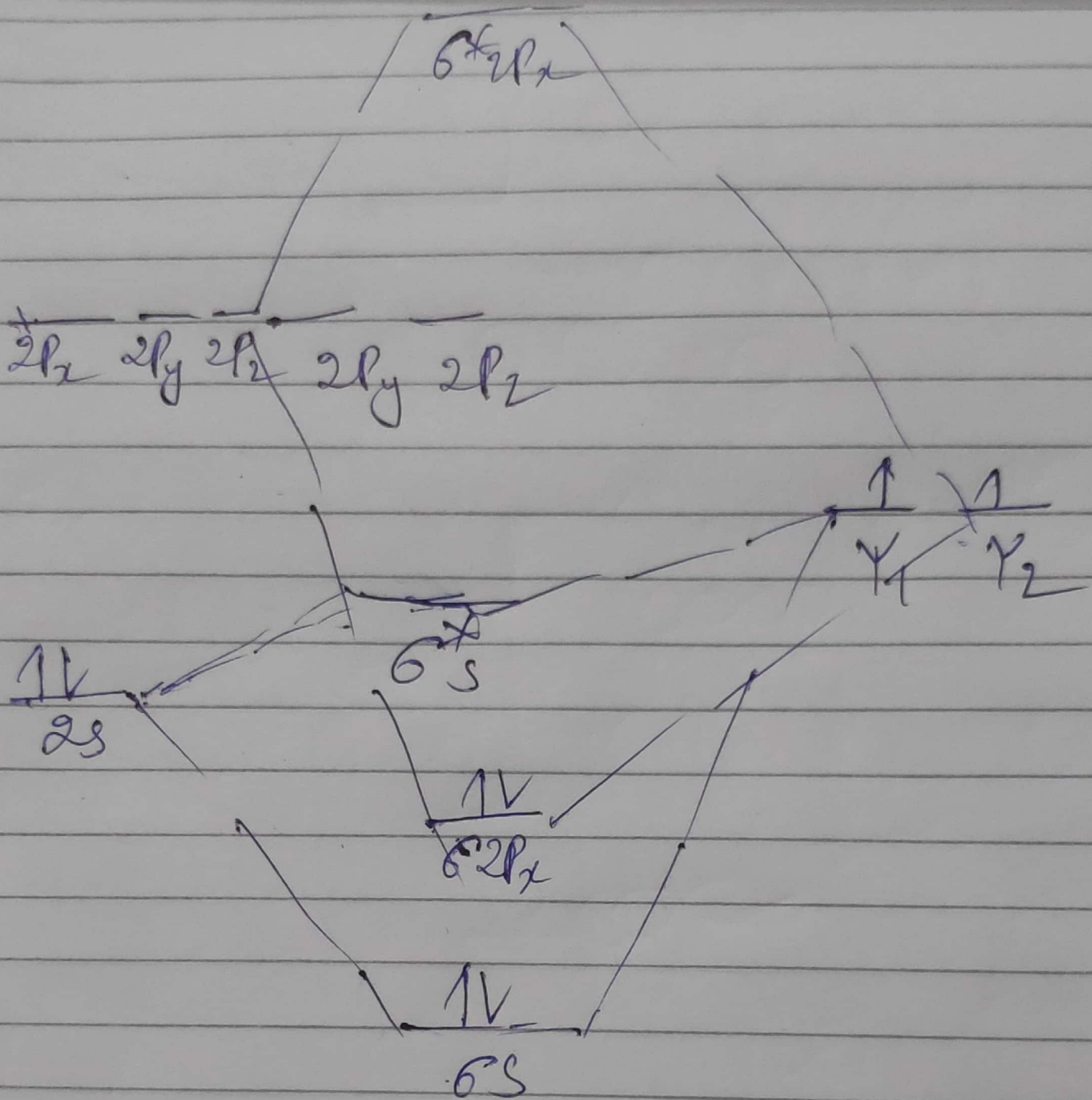


$$\psi_1 = 1s_{H_1} + 1s_{H_2}$$

$$\psi_2 = 1s_{H_1} - 1s_{H_2}$$

Now these group orbital overlap with orbital of Be to form four molecular orbital as given below:-





orbital of
Be

M.O of
BeH₂

group orbitals
of H